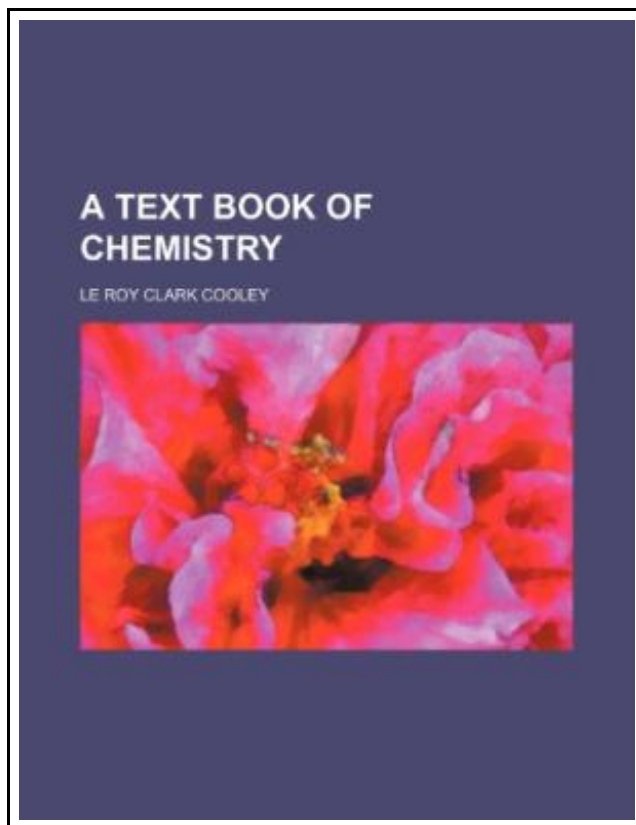


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1869 Excerpt: .gaseous. In the colors of these gases we notice a curious gradation. Bromine is at one- end of the spectrum, dark red; chlorine in the middle part, greenish yellow; and iodine at the other end, a beautiful violet. 2. Their chemieal properties.--These four elements resemble each other in their chemical properties also. They combine with the same substances, and generally in the same proportions. For hydrogen and the metals they all have strong attraction; with oxygen they (with the possible exception of fluorine) unite with a feeble force; with nitrogen they form explosive compounds. They are all univalent, and the single compound which each one forms with hydrogen is an acid. The following symbols illustrate the analogous composition of the compounds of these elements: 1. What names shall be given the compounds of hydrogen with these elements? 2. What number of atoms of each element in these acids? 3. What proportions by volume? 4. What proportions by weight? The compounds of these elements with oxygen, whose symbols are given, are not all known in a free state; most of them are found combined with water, and in this case they are acids. The Cl_2O_7 is then changed to HClO_4 , which is called perchloric acid. $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} = \text{H}_2\text{ClO}_4$ 1. Name the other acids corresponding to the several oxides represented. 2. What would be the symbol for each? 3. Describe the composition of each acid:...



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